

## Cooperative Project II: Identification, Properties and Synthesis of an Unknown Ionic Compound

For cooperative project II, we will follow the lab on page 115 of the *Cooperative Chemistry Laboratory Manual, 5<sup>th</sup> edition*. This lab will require each group to: 1. Investigate the chemical hazards of some of the chemicals used in this lab; 2. identify an unknown ionic compound; 3. investigate the chemical and physical properties of the unknown ionic compound; 4. devise two syntheses of the compound and compare them for cost effectiveness, safety and potential yield. This aspect of the project will not be carried out in lab, but will be included in your weekly report.

Before coming to lab, be sure to read all background reading assignments, complete all practice problems and any other prelab assignment required by your instructor.

### Requirements:

- Consult the appropriate material safety data sheet (MSDS) and find the LD<sub>50</sub> or LC<sub>50</sub> value for hydrochloric acid, sulfuric acid and acetic acid and include these values in your lab report.
- Be sure to read the lab on page 115 of the *Cooperative Chemistry Laboratory Manual, 5<sup>th</sup> edition*.
- Identify the unknown compound. You will only be given 5 grams of the unknown compound.
- Discover as many chemical and physical properties of the compound as you can.
- Devise two syntheses of the compound and compare them for cost effectiveness, safety and potential yield. Be sure to reference all data you obtain from outside sources. The safety analysis must include the LD<sub>50</sub> for each of the reactants and products in each of your devised reactions.
- Below are guidelines for referencing data from different scientific sources.

### Reference Formats:

Reference format for referencing a journal article:

- Author, A. A.; Author, B. B.; Author, C. C. Title of Article. *Journal Abbreviation (italics)* [Online if online] **Year (boldface)**, *Volume (italics)*, Pageination.
- Example:
  - Slunt, K. M.; Giancarlo, L. C. Student-Centered Learning: A Comparison of Two Different Methods of Instruction. *J. Chem. Educ.* **2004**, *81*, pp 985-988.
- Reference format for referencing a scientific handbook:
  - Editor, A. A., Editor, B. B., Eds. *Handbook Title (italics)*, Edition number [Online if online]; Publisher: Place of Publication, Year; Pageination or other identifying information.
  - Example:
    - Lide, D. R., Ed. *CRC Handbook of Chemistry and Physics*, 84th ed. [Online]; CRC Press: Boca Raton, FL, 2003; p 83.
- Reference format for referencing an MSDS:
  - **Hard copy (paper) MSDS**
    - *Titanium Dioxide*; MSDS No. T3627; Mallinckrodt Baker: Phillipsburg, NJ, November 12, 2003.
    - If there is not a MSDS No., *Titanium Dioxide*; MSDS; Mallinckrodt Baker: Phillipsburg, NJ, November 12, 2003.
  - **MSDS obtained from an Internet search**
    - *Titanium Dioxide*; MSDS No. T3627; Mallinckrodt Baker: Phillipsburg, NJ, November 12, 2003. <http://www.jtbaker.com/msds/englishhtml/t3627.htm> (accessed 4/15/04).
    - If there is not a MSDS No., *Titanium Dioxide*; MSDS; Mallinckrodt Baker: Phillipsburg, NJ, November 12, 2003. <http://www.jtbaker.com/msds/englishhtml/t3627.htm> (accessed 4/15/04).

### Requirements for keeping a laboratory notebook:

- Each member of your group is required to keep a scientific notebook. Your notebook is where you will record your day-to-day activities in the lab. You should describe experiments as you do them and note observations as you make them. Record and analyze your data in your notebook.
- The notebook must be a carbonless, duplicate page notebook.
- Read "Recording and Reporting Results", "The Laboratory Notebook" on page 17 of the laboratory manual.
- Be sure to keep the following in mind when recording data in your notebook:
  - Include a descriptive "Table of Contents"
  - Pages must be numbered
  - Include date data was collected
  - Record raw data
  - Record procedures and observations
  - Write in ink
  - Mistakes must be crossed out appropriately (single line through the mistake with your initials)

### Requirements for writing weekly summaries:

- List all group members present for the lab
- Give a short overview of the lab
- Present results in tables and/or graphs
- Discuss future work

### Safety Precautions:

- **Bunsen Burners:** The Bunsen burner is a common device used for heating. It is limited in usefulness to warming aqueous solutions or other nonflammable substances. Care must be used when using a Bunsen burner. Your laboratory instructor will demonstrate the proper method for adjusting, lighting, and using a Bunsen burner.
- Wear your safety goggles at all times. You will be using several different acid, base and salt solutions. If you get any solution on you, wash immediately with lots of water.
- Dispose of all waste in the labeled hazardous waste container in the common equipment area. Use a wash bottle to rinse glassware into the container.
- You will use several different reagents throughout this experiment. Many of the reagents are located in the common reagent area in the front of the lab. When you need a reagent, take it to your lab bench, use it, put the lid back on it and then put it right back. Do not carry open bottles of solutions to the reagent bench (poor etiquette). Be sure you return it to its proper place.
- Never pour anything back into the stock solution.
- Never lay reagent bottle stoppers or caps on the lab bench. The entire reagent may become contaminated. In addition, the residue on the bench may be hazardous and linger for days or weeks. This could injure someone well after the fact. Hold the stopper in your other hand while you get the material out of the bottle. Replace stoppers immediately and completely.
- **Centrifuge safety:** Do not open the centrifuge until it has completely stopped. Do not try to manually stop the centrifuge from spinning. The blank tube filled with water is necessary to balance the instrument. Do not walk away from a spinning centrifuge.
- Balances are especially sensitive, expensive devices. **Never weigh chemicals directly on the pan.** Use weighing paper or boat or a container such as a beaker or flask. Remove the container from the balance, add the chemical and then replace the container. If you spill anything onto the balance, please notify the instructor **immediately**.

## Background Reading and Practice Problems:

### *Cooperative Chemistry Laboratory Manual*

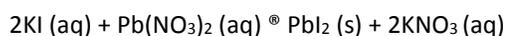
- Safety Rules, Laboratory Etiquette (12-14)
- NFPA Hazard Codes, Waste, MSDS's (15-16)
- Recording and Reporting Results (17-19)
- Writing Lab Reports (19-23)
- Reporting Numerical Results, Significant Figures, Graphs (35-40)
- Measuring Devices (47-49)
- To Make up a Solution of Known Concentration (64)
- Bunsen Burner (53)
- Laboratory Techniques (55-63)
- Filtration (70)
- Identification, Properties and Synthesis of an Unknown Ionic Compound (p. 115)

### *General Chemistry, 4<sup>th</sup> Edition by McQuarrie*

- Ions, Ionic Compounds: Read sections 2.7, 2.12, 6.1-6.6
- Reactions: Read Section 10.2 and 10-9; Complete Practice Problems 10-13, 10-14, 10-15, and End of Chapter Problems 10-52

**Additional Background Information:** Chemical analysis can be divided according to quantitative or qualitative analysis. In the last experiment, we were asked to quantitatively find the density of various materials. However, sometimes a chemist is interested in what species are present in a sample, as opposed to the amount of species present in a sample. Thus, a qualitative analysis involves the determination of the types (not amounts) of ions present in a solution. A "qualitative analysis scheme" is then a systematic procedure for separating and identifying various ions present in an aqueous solution.

- The procedures in most qualitative analysis experiments involve the use of using **precipitation reactions**. A precipitation reaction is a reaction in which an insoluble solid compound is formed when solutions of two soluble compounds are mixed. The solid compound that forms during a precipitation reaction is referred to as the **precipitate**. An example of a precipitation reaction is the reaction between aqueous solutions of KI and  $\text{Pb}(\text{NO}_3)_2$ :



- We know that the ionic compounds KI,  $\text{Pb}(\text{NO}_3)_2$ , and  $\text{KNO}_3$  are soluble in water and  $\text{PbI}_2$  is not soluble in water, see "General Solubility Guidelines of Ionic Compounds in Water" on page 60 of the *Cooperative Chemistry Laboratory Manual*.
- There are some general procedures in qualitative analysis that will be used throughout this experiment. One is the use of a centrifuge to aid the separation of a precipitate from a solution in a test tube. There is a centrifuge at the end of each lab bench. To use the centrifuge, make sure the test tube containing the precipitate and solution (sample test tube) is not overly full. Place the sample test tube in one of the centrifuge tubes. Place a blank tube containing the same amount of water as you have in your sample test tube in the centrifuge tube that is opposite your sample test tube. Turn on the centrifuge and allow it spin for about 1 minute. Once you have centrifuged your sample test tube, now you can **decant** the supernatant liquid from the precipitate. Simply pour off (decant) the liquid that is above the precipitate (the supernatant).

### Qualitative Procedure for the Analysis of $\text{Ca}^{2+}$ and $\text{Mg}^{2+}$ Ions

1. Obtain 1 mL (approximately 20 drops) of the solution to be analyzed in a small test tube.
  2. Add 2 drops of 6M  $\text{HNO}_3$  and then add 1 mL of 6M  $\text{NH}_4\text{OH}$ .
  3. Add 1 mL of 0.3 M  $(\text{NH}_4)_2\text{CO}_3$  and stir. If a precipitate forms,  **$\text{Ca}^{+2}$  ions are present**. Centrifuge and decant the supernatant, which may contain  $\text{Mg}^{+2}$ . Save this solution for analysis in step 4. If no precipitate forms upon the addition of the 0.3 M  $(\text{NH}_4)_2\text{CO}_3$ ,  $\text{Ca}^{2+}$  are not present and you should continue with step 4.
  4. To the solution from step 3, add 1mL of  $\text{K}_2\text{HPO}_4$  solution, stir and let stand 1 minute. The slow formation of a **white** precipitate indicates **magnesium**.
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