

**Problem Set 2 – Due by 5 PM Friday, Sept. 20, 2019****Part A**

Please give the compound formula corresponding to each name below

	Compound name	Compound formula
1.	iodine heptafluoride	<b><i>IF<sub>7</sub></i></b>
2.	iron(III) dichromate	<b><i>Fe<sub>2</sub>(Cr<sub>2</sub>O<sub>7</sub>)<sub>3</sub></i></b>
3.	calcium nitride	<b><i>Ca<sub>3</sub>N<sub>2</sub></i></b>
4.	lithium carbonate	<b><i>Li<sub>2</sub>CO<sub>3</sub></i></b>
5.	tetraphosphorus decoxide	<b><i>P<sub>4</sub>O<sub>10</sub></i></b>
6.	xenon trioxide	<b><i>XeO<sub>3</sub></i></b>
7.	barium perchlorate	<b><i>Ba(ClO<sub>4</sub>)<sub>2</sub></i></b>
8.	sulfur tetrachloride	<b><i>SCl<sub>4</sub></i></b>
9.	potassium cyanide	<b><i>KCN</i></b>
10.	dinitrogen monoxide	<b><i>N<sub>2</sub>O</i></b>

**Part B**

Please give the compound name corresponding to each of the following formulas.

	Compound formula	Compound name
11.	SeCl <sub>6</sub>	<b><i>selenium hexachloride</i></b>
12.	Pd(CH <sub>3</sub> COO) <sub>2</sub>	<b><i>palladium(II) acetate</i></b>
13.	K <sub>2</sub> CrO <sub>4</sub>	<b><i>potassium chromate</i></b>
14.	KrF <sub>2</sub>	<b><i>krypton difluoride</i></b>
15.	Ga <sub>2</sub> S <sub>3</sub>	<b><i>gallium sulfide</i></b>
16.	Mn <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	<b><i>manganese(II) phosphate</i></b>
17.	P <sub>2</sub> I <sub>4</sub>	<b><i>diphosphorus tetraiodide</i></b>
18.	SbBr <sub>5</sub>	<b><i>antimony pentabromide</i></b>
19.	WO <sub>3</sub>	<b><i>tungsten(VI) oxide</i></b>
20.	N <sub>2</sub> F <sub>2</sub>	<b><i>dinitrogen difluoride</i></b>