

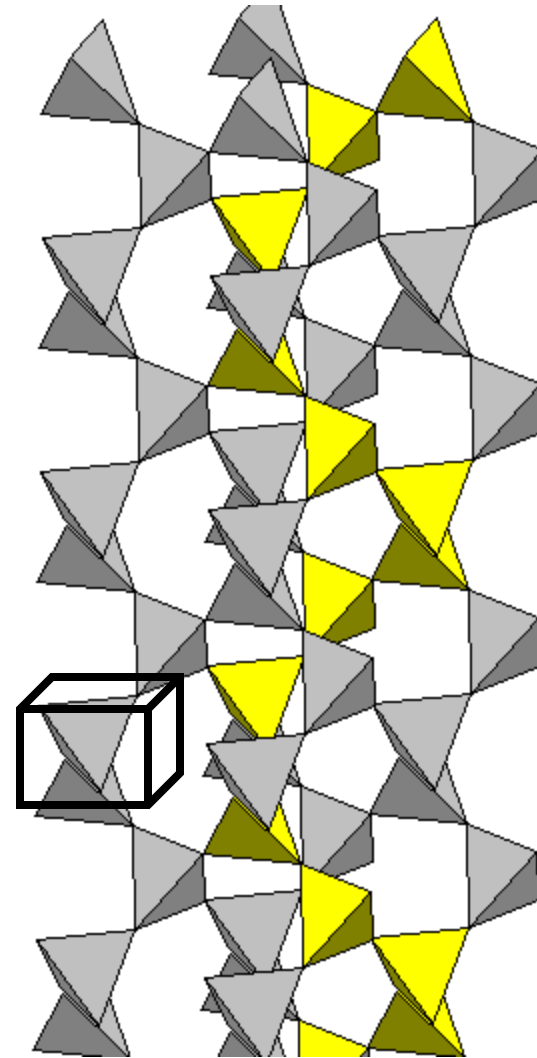
A Short History of the Universe

- ◆ The universe in a grain of sand
- ◆ Particles and forces
- ◆ The Big Bang
- ◆ Early history of the universe
- ◆ Nucleosynthesis

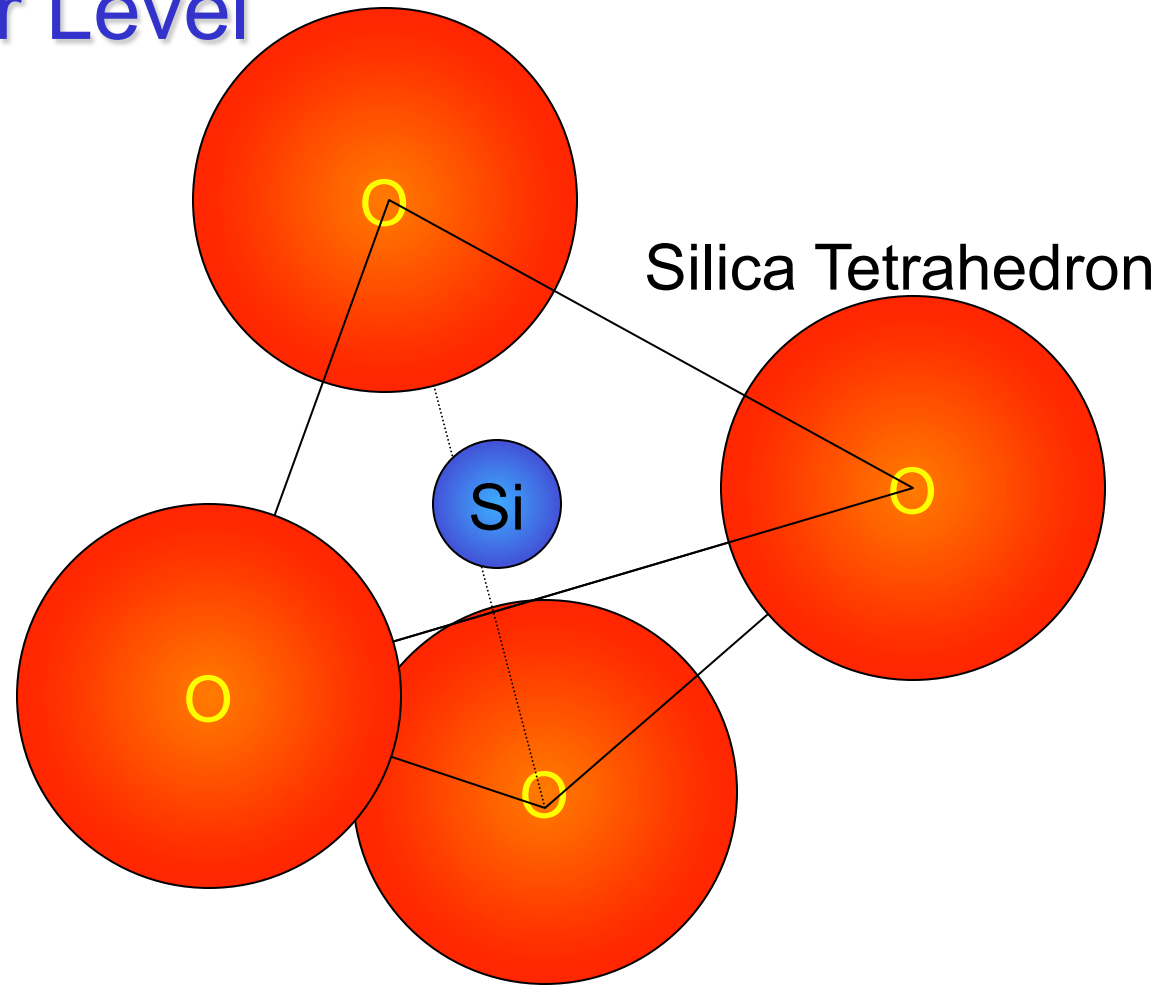
Molecular Level

Quartz – SiO_2

Composed of linked
spirals of silica
tetrahedrons



Molecular Level



Each tetrahedron is made of one silicon atom bonded to four oxygen atoms

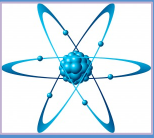
The atom is the fundamental particle of chemistry. It is composed of three smaller types of particles

e⁻

Electrons are negatively charged particles with very little mass that are attracted to the positively charged nucleus of the atom



Protons (positively charged particles) and **neutrons** (particles with no charge) are bound together in the relatively massive nucleus of the atom.

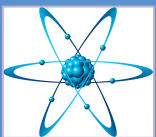


All hydrogen atoms have one proton

All nickel atoms have twenty-eight protons

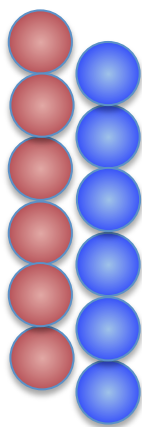
H																			He
Li	Be											B	C	N	O	F		Ne	
Na	Mg											Al	Si	P	S	Cl		Ar	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br		Kr	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I		Xe	
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At		Rn	
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub								
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu			
		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr			

An element is all atoms with the same number of **protons** in the nucleus.



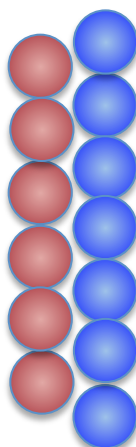
The number of **neutrons** in the nucleus can vary from atom to atom in an element. These variations, called isotopes, have different atomic weights, and some are not stable, and thus can radioactively decay. The isotopes are chemically identical.

All carbon atoms have six **protons**



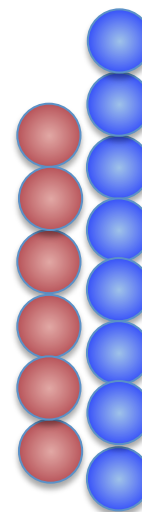
Carbon-12 has 6
neutrons

Stable



Carbon-13 has 7
neutrons

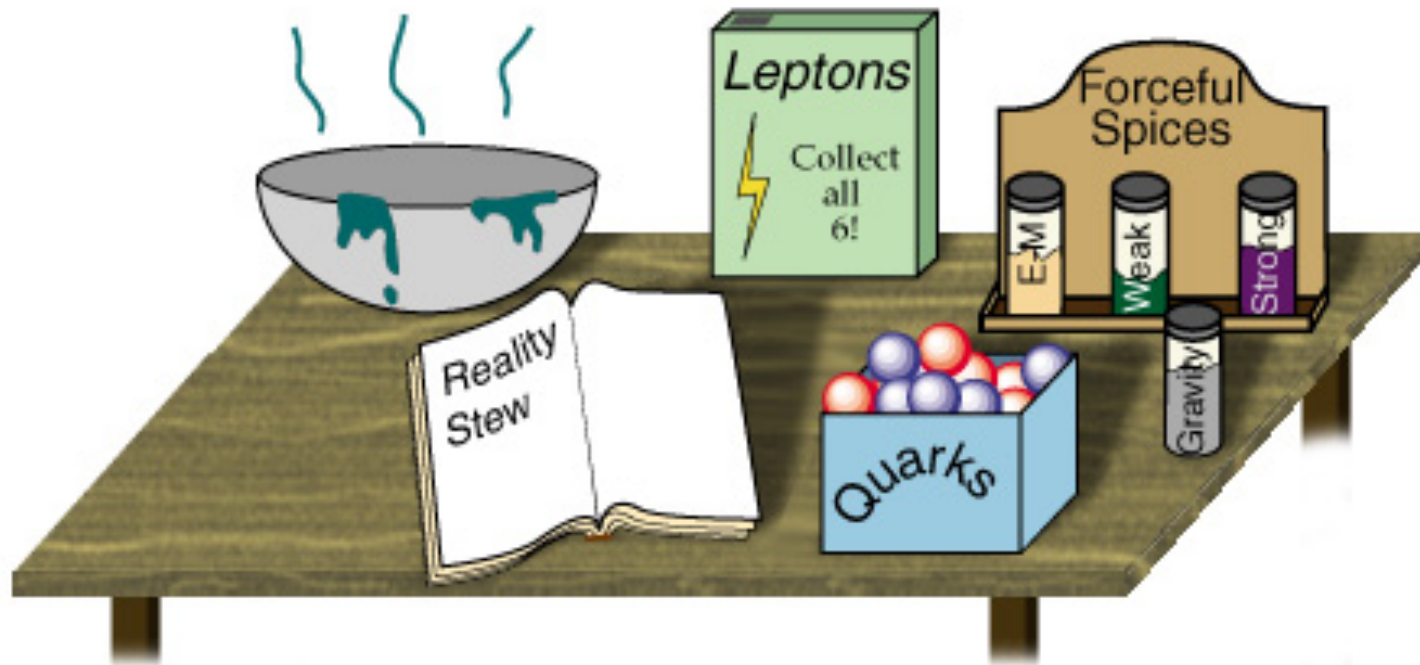
Stable



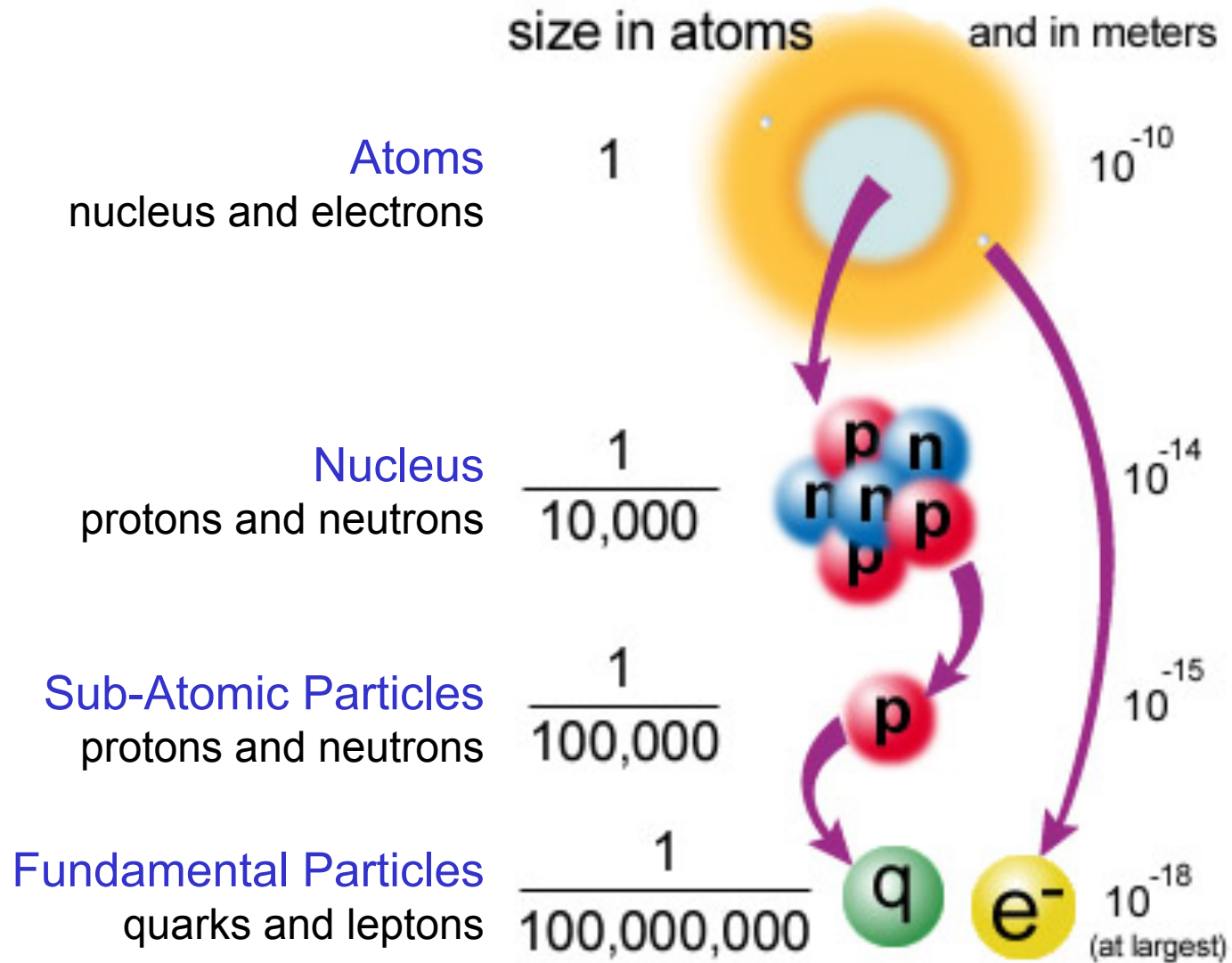
Carbon-14 has 8
neutrons

UNSTABLE!

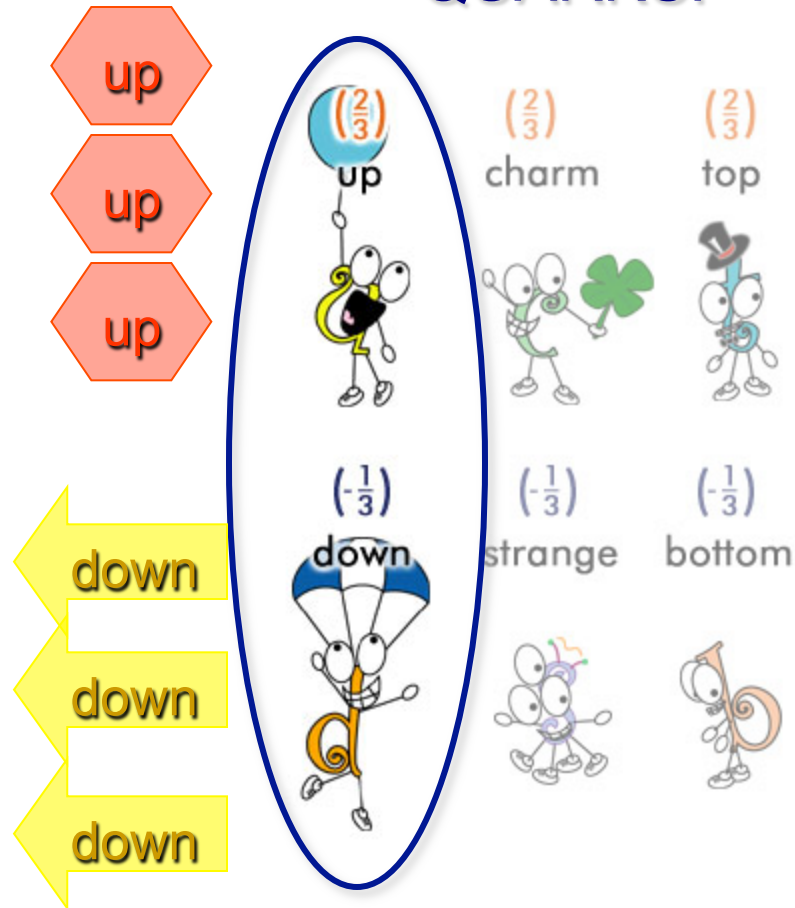
Sub-Atomic Level



Protons and neutrons are made up of *fundamental particles* called quarks. Electrons are fundamental particles.

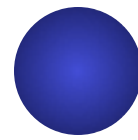


QUARKS!



Fundamental Particles

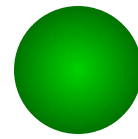
Each quark has a fractional electrical charge. They combine to form sub-atomic particles



Proton

up + *up* + *down*

$$(+2/3) + (+2/3) + (-1/3) = +1 \text{ charge}$$

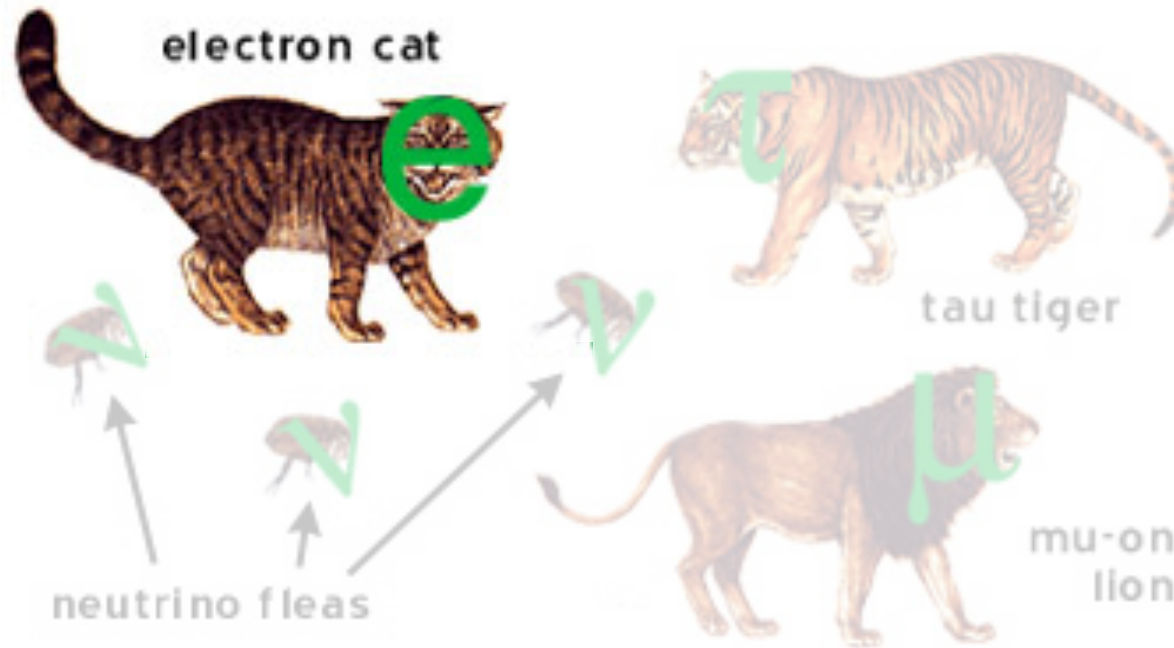


Neutron

up + *down* + *down*

$$(+2/3) + (-1/3) + (-1/3) = \text{no charge}$$

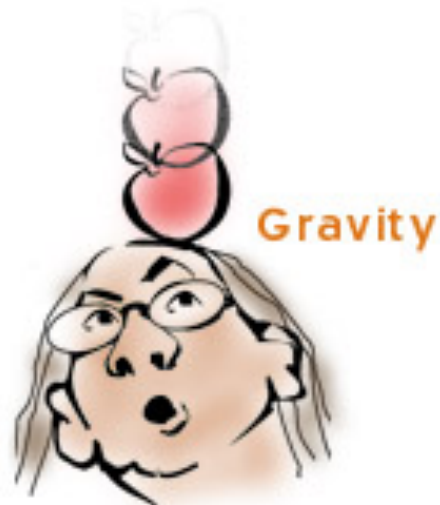
Fundamental Particles



LEPTONS!

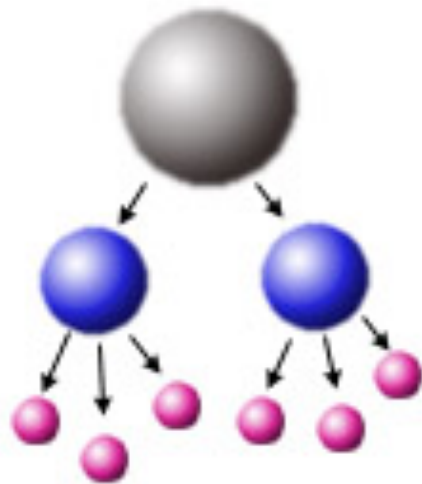
Electrons are the most important lepton.

Forces



Weak Nuclear Force

Responsible for holding more massive quarks and leptons together



No further decay
is possible!

More massive fundamental particles rapidly decay to more stable, lower mass particles (e.g., tau to muon to electron)

Strong Nuclear Force

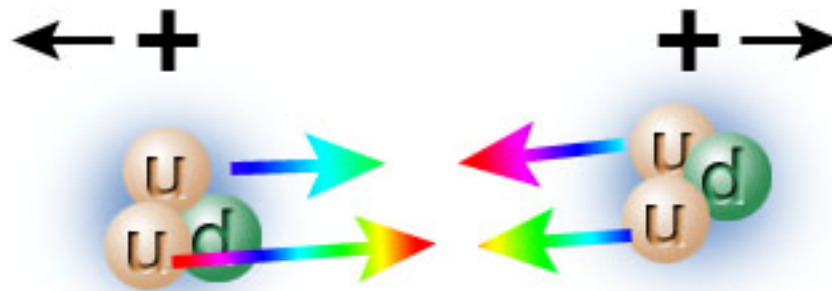
Holds quarks together in sub-atomic particles



Attraction is caused by differences in
“quark color” (which has nothing to do
with real color)

Residual Strong Nuclear Force

Holds nucleus together

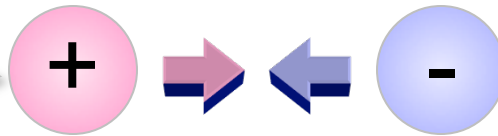


“Color” attraction between quarks
overcomes electromagnetic repulsion, but
only over very small distances

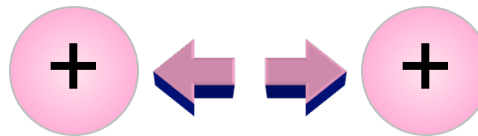
Electromagnetism

attraction of oppositely charged particles

e.g., a proton (+1 charge) and an electron (-1 charge)



repulsion of similarly charged particles

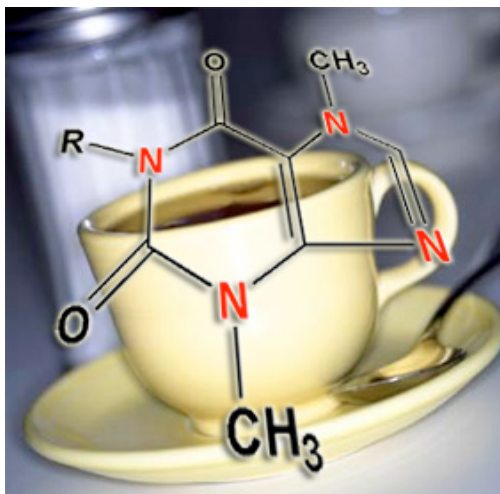


e.g., two protons (each with a +1 charge)

Electromagnetism

Atoms bond together to make molecules, which are described with chemical formulas. Molecular structures can be simple, or very complex.

Caffeine



1,3,7-trimethyl-1H-purine-2,6(3H,7H)-dione

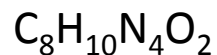
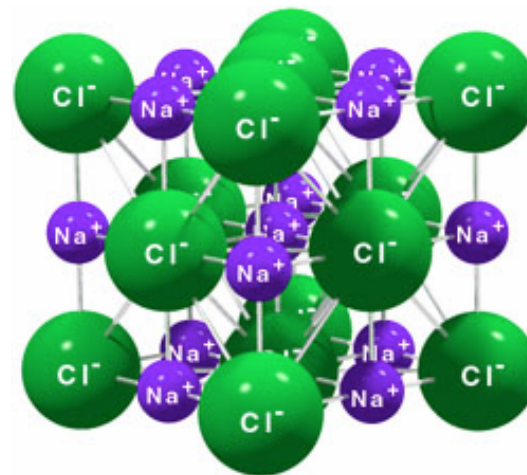


Table Salt



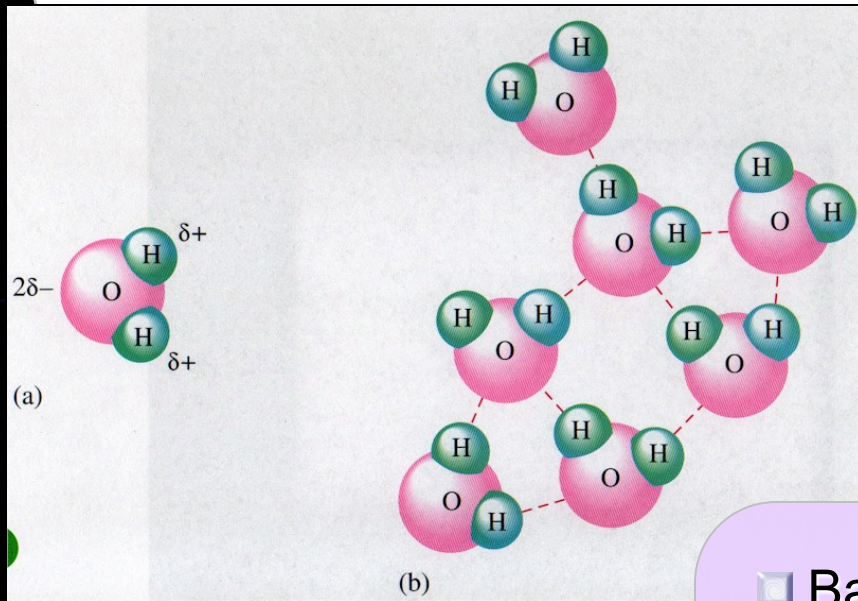
Sodium Chloride



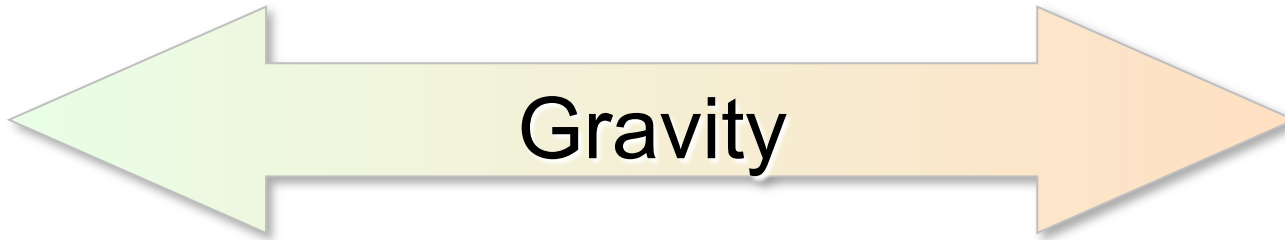
<http://www.mylot.com/w/image/2435174.aspx>

<http://www.chemistry.wustl.edu/~edudev/LabTutorials/Water/PublicWaterSupply/PublicWaterSupply.html>

Electromagnetism



- Basis of all chemistry
- Responsible for properties of phases of matter
- Prevents everyone in this room from falling to the center of the Earth



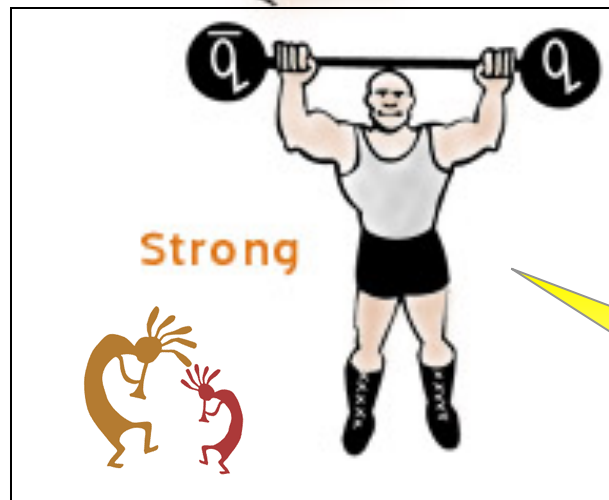
Gravity is a property of matter – any object with mass also exerts gravitational attraction.

All objects attract each other gravitationally, with more massive objects exerting greater force.

It has been hypothesized that mass warps space in a higher dimension, thus causing objects to be attracted to each other.

Forces

mutual attraction of all matter slows universal expansion



bonds atoms to each other, responsible for electric and magnetic fields

is not important in current universe

holds subatomic particles *AND* atomic nuclei together.

So how did the universe get to be this way?

Element Formation in Stars

Stars and Galaxies

Start Point
Elements: H, He, Li

The Big Bang

Planetary System Formation

Forming Earth-like Planets

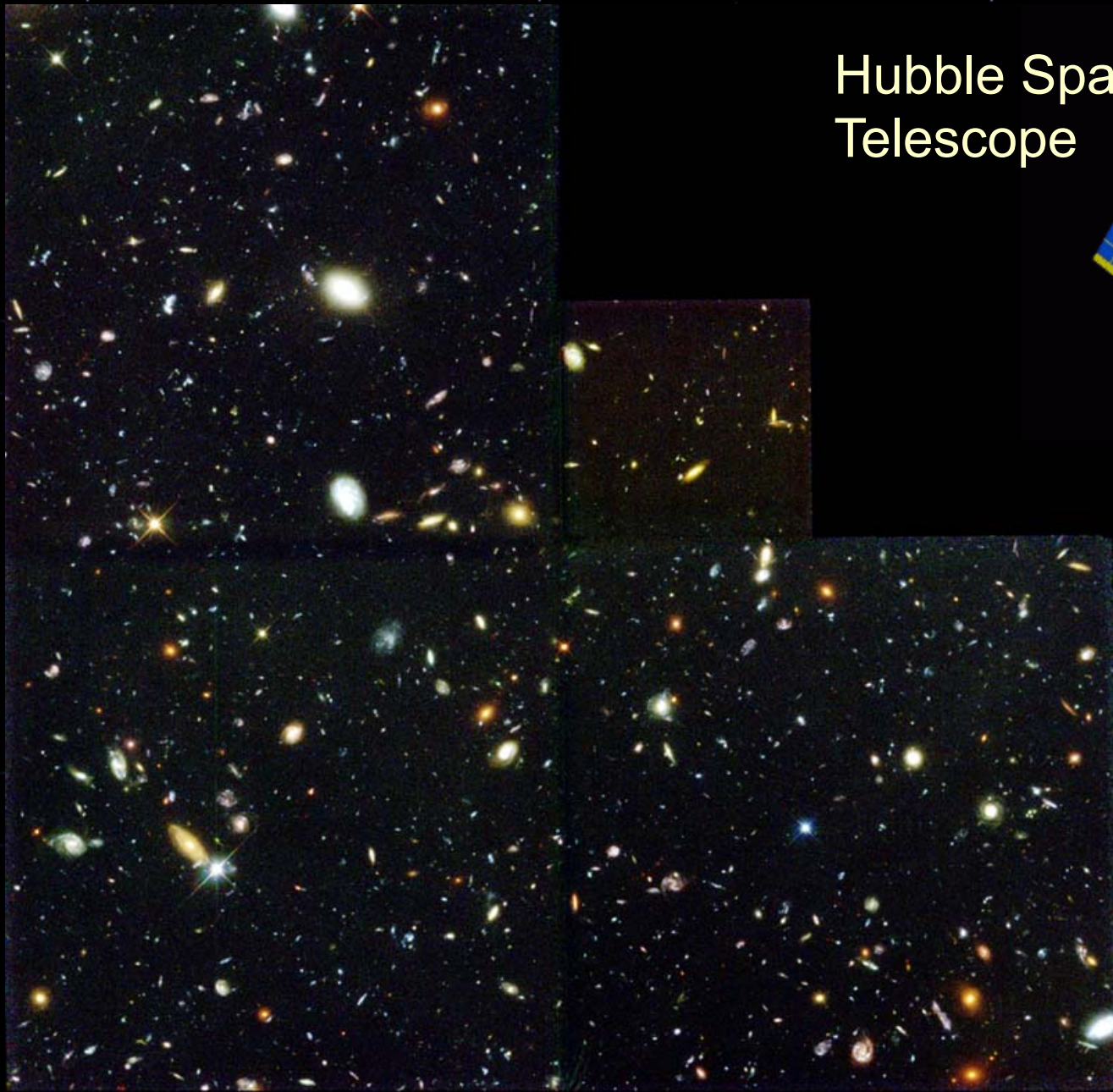
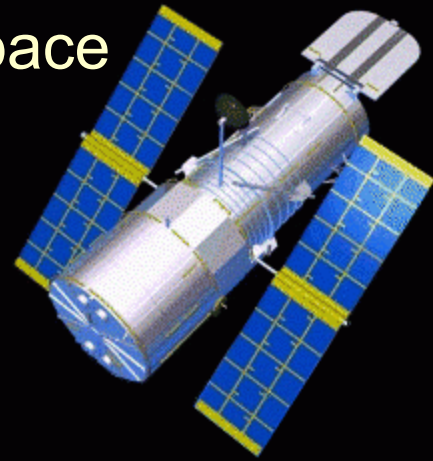
Our Solar System

Forming Jupiter-like Planets

Chemistry of Life



Hubble Space
Telescope



Mapping to Date:
Millions of galaxies

Hubble Deep Field

HST WFPC2

ST ScI OPO January 15, 1996 R. Williams and the HDF Team (ST ScI) and NASA

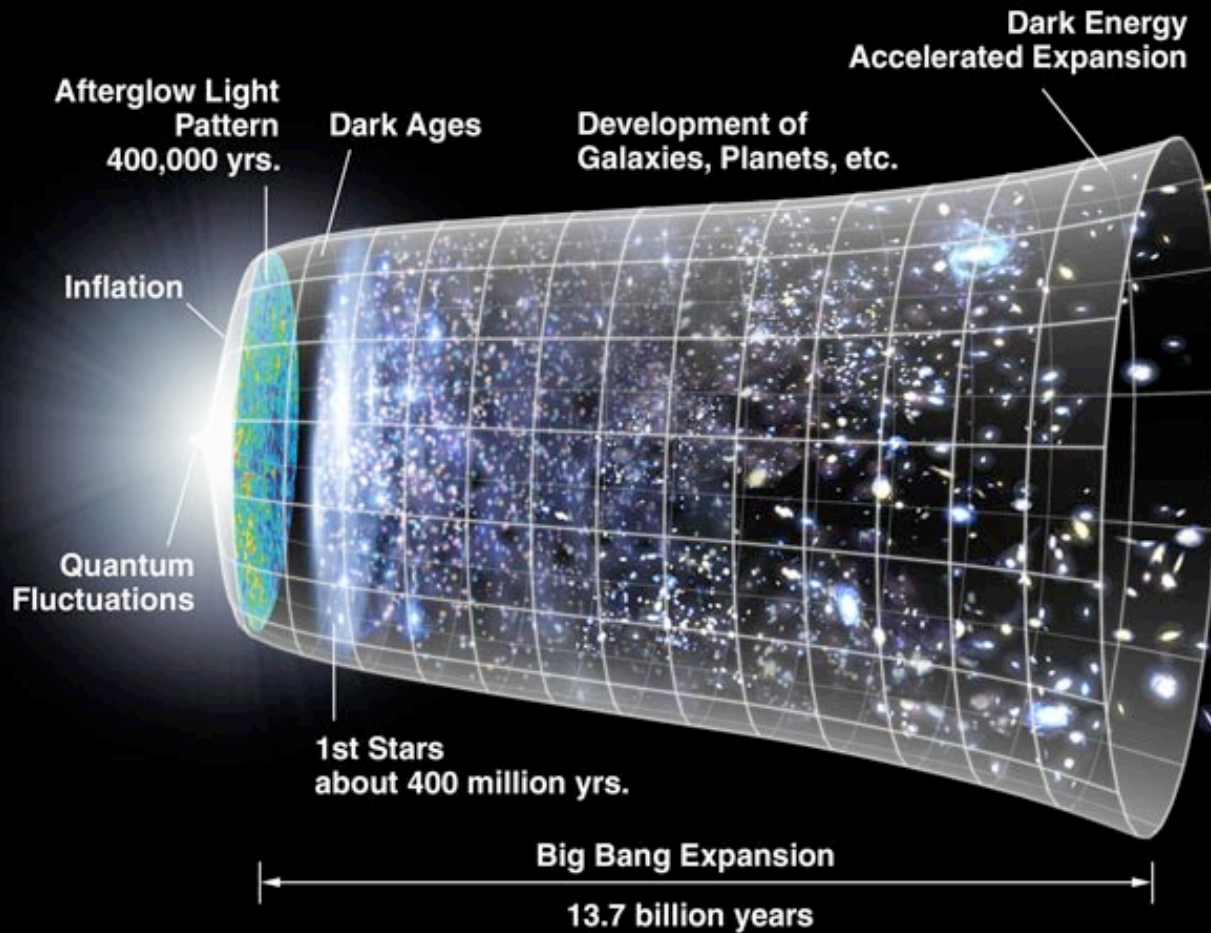
● Humans have observed only a small portion of the Universe, but other than our members of our own galaxy and our galaxy's nearest neighbors, all objects we have observed are moving away from us.

● There's nothing wrong with us - the objects are moving away from each other, too. In other words, the universe appears to be expanding.

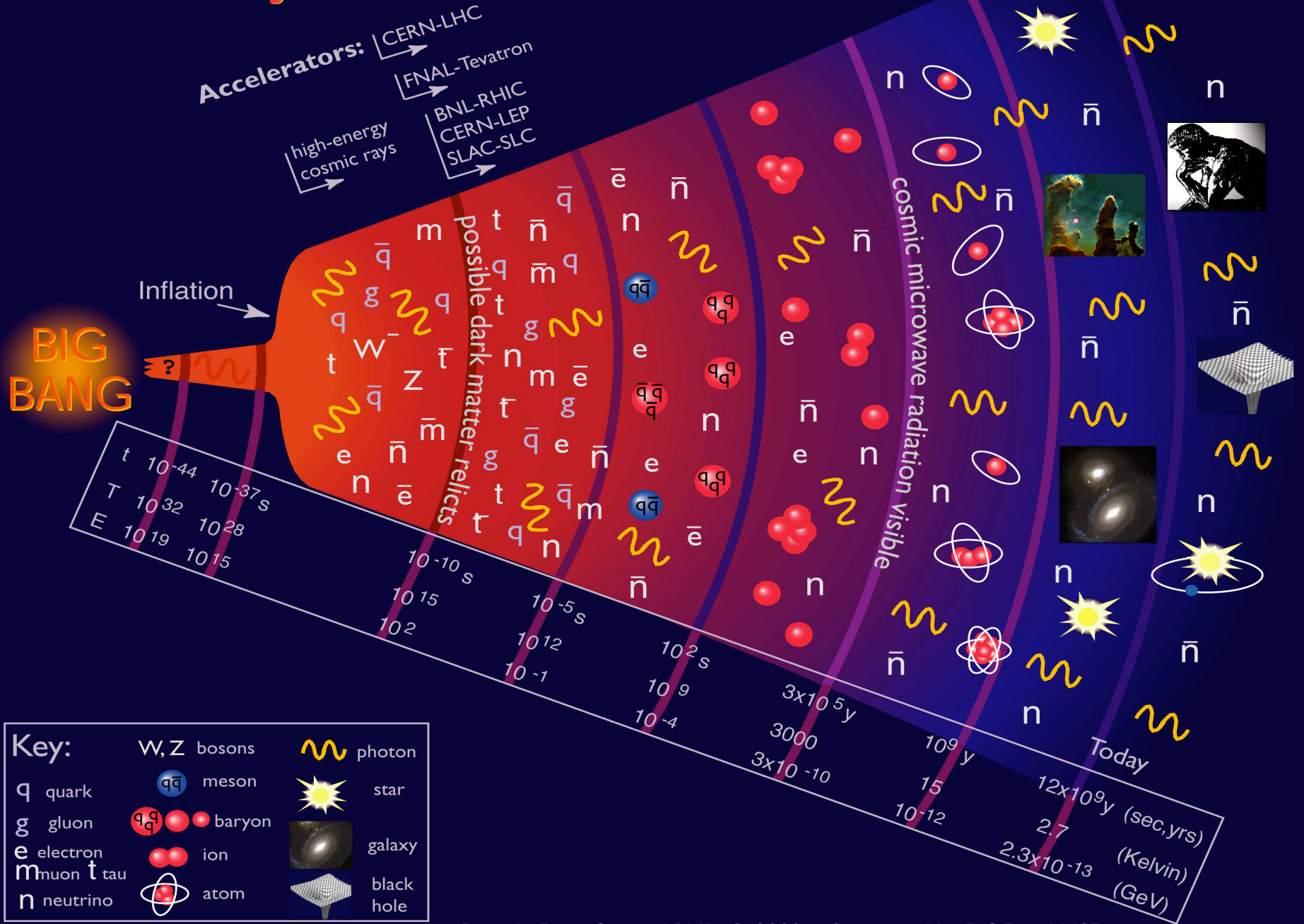
● If we run the tape backwards, and re-compress the universe, there was a time (12-14 billion years ago), when the entire universe was unimaginably dense and small.

● The event that led to the expansion from that dense universe is called the "Big Bang."

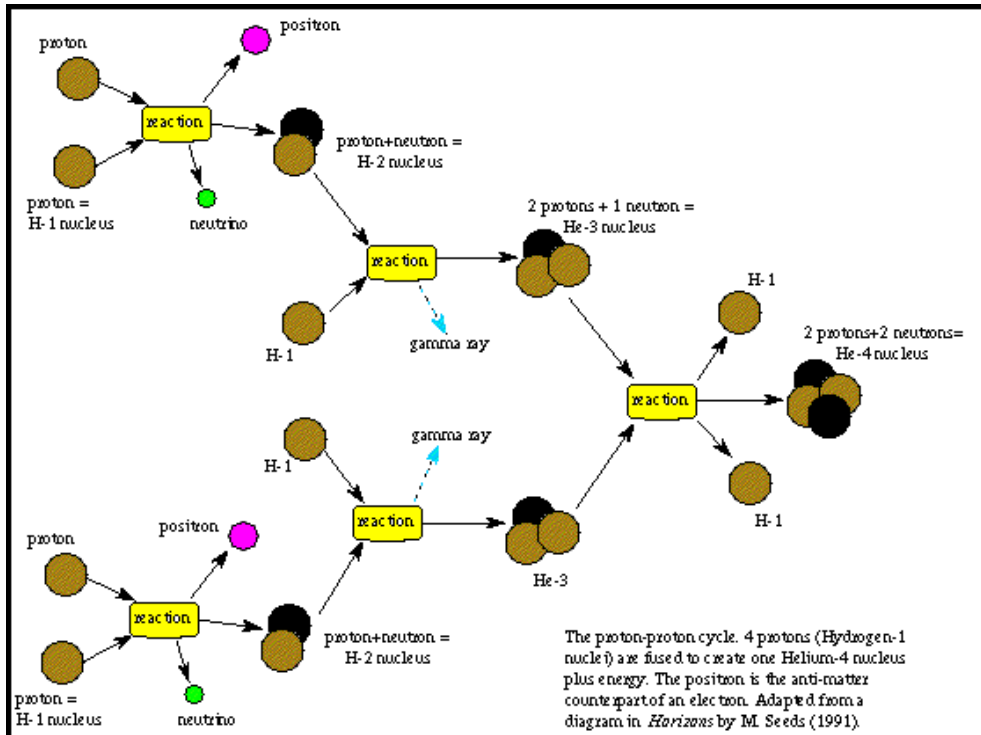
The leading hypothesis in cosmology is that the universe today is the result of expansion after the Big Bang.



History of the Universe



Nucleosynthesis – Making New Elements

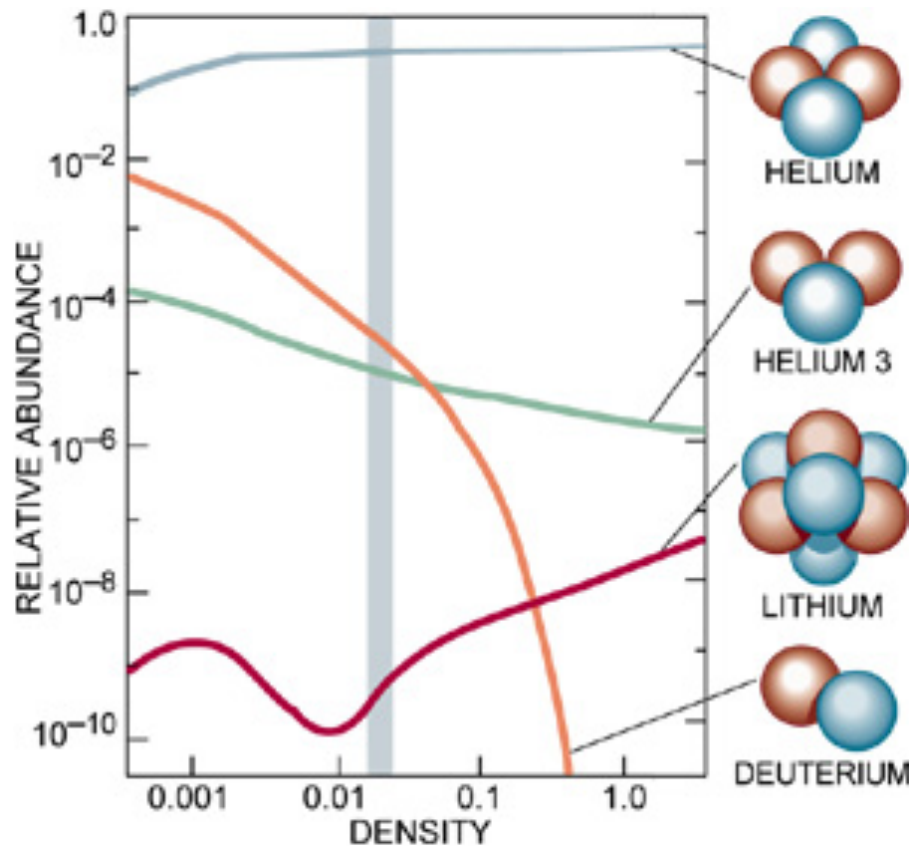


Hydrogen Fusion

Given enough energy, protons can “hit” each other hard enough for the strong nuclear force to overcome electromagnetic repulsion.

In the modern universe, hydrogen fusion takes place in the hearts of stars. It is also the basis of thermonuclear bombs.

Big Bang Elements



Baryons (protons and neutrons) formed at a time when there was still lots of energy available to shoot them around.

As the baryons knocked into each other, some stuck, forming isotopes of the elements hydrogen, helium and lithium.

There was not enough energy to form heavier elements.

Big Bang Elements

Hydrogen (1)								Helium (2)										
H																		He
Li	Be											B	C	N	O	F		Ne
Na	Mg											Al	Si	P	S	Cl		Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br		Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I		Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At		Rn
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub							
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu		
		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		

Elements formed during and soon after the Big Bang.

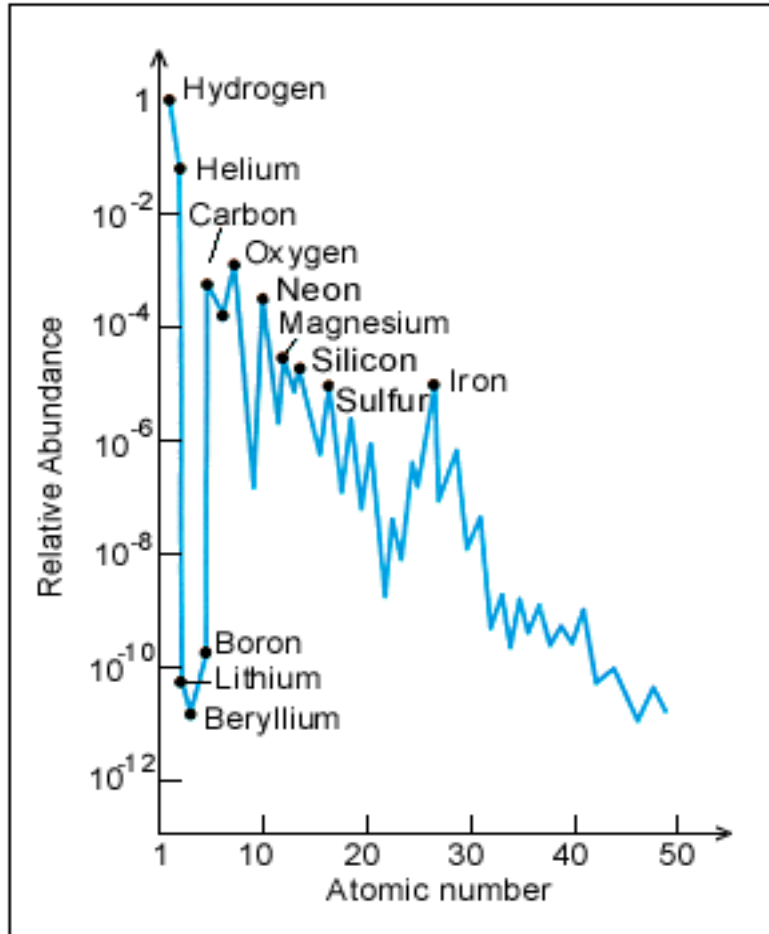
Elements of the Earth's Crust

Noble Gasses

H																			
Li	Be											B	C	N	O	F			
Na	Mg											Al	Si	P	S	Cl			
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br			
Rb	Sr	Y	Zr	Nb	Mo		Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I			
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi					
		La	Ce	Pr	Nd		Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu			
			Th		U														

Unstable Elements

Composition of Star Nebulae



H	1,000,000
He	98,000
O	850
Ne	120
N	100
Fe	47
Na	2
Other	100

The modern, visible universe is still primarily hydrogen and helium, but with other elements

These elements were not made during the Big Bang, and must have been assembled in the billions of years since then.

Pencil Nebula • NGC 2736



Hubble
Heritage

NASA and The Hubble Heritage Team (STScI/AURA) • Hubble Space Telescope ACS • STScI-PRC03-16

Post-Big Bang Elements

Where were (and are!)
these heavier elements
formed?

Both in the hearts of living
stars and during some of
their explosive deaths...